

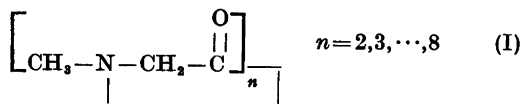
# Crystal Structure of Cycloheptasarcosyl Hydrate

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The crystals belong to the monoclinic system with space group  $P2_1/c$  and cell dimensions  $a = 28.725(5) \text{ \AA}$ ,  $b = 11.048(1) \text{ \AA}$ ,  $c = 11.036(1) \text{ \AA}$ ,  $\beta = 132.52(1)^\circ$ . The phase problem was solved by direct methods, and the  $R$ -value arrived at for 2250 observed reflections was 6.9% ( $R_w = 4.9\%$ ). The conformation is *cis,cis,cis,cis,trans,trans,trans*. Water molecules form hydrogen bond bridges linking molecules to endless chains along [001]. Bond distances and angles are compared with those of cyclopentasarcosyl, cyclooctasarcosyl, and cycloalanyl-tetrasarcosyl.

Cyclic oligopeptides of sarcosine of the general formula (I) have been studied by Dale and Titlestad.<sup>1-3</sup> To account for the relatively high observed resistance to ring inversion,



transannular interactions between N and C (carbonyl) were suggested. Recent findings<sup>4-6</sup> do not support this assumption, and the explanation must be sought in the intrinsic conformation of the peptide chain itself.<sup>2</sup> For the case  $n = 7$  the NMR-spectrum is complex with broad methylene and *N*-methyl lines which do not become simpler on cooling to  $-70^\circ\text{C}$ . Benzene addition resolves seven main *N*-methyl signals together with several less intense signals, indicating that one major conformation is present. The crystal conformation which is obtained by dissolution in  $\text{CH}_2\text{Cl}_2$  at  $-80^\circ\text{C}$ , shows three *N*-methyl lines (intensity 9:9:3) and seven partially resolved quartets. In order to settle the conformational problem, and to obtain detailed information of the molecular geometry, and X-ray crystallographic investigation of cycloheptasarcosyl has been carried out.

The crystals belong to the monoclinic system and the systematic absences lead to the space group  $P2_1/c$ . The cell parameters measured by means of a four circle diffractometer, and their estimated standard deviations are:

$$a = 28.725(5) \text{ \AA}, b = 11.048(1) \text{ \AA}, c = 11.036(1) \text{ \AA}, \beta = 132.52(1)^\circ$$

The unit cell contains four  $\text{C}_{21}\text{O}_7\text{N}_7\text{H}_{35}\cdot\text{H}_2\text{O}$  formula units.

With  $2\theta_{\text{max}} = 50^\circ$  and  $\text{MoK}\alpha$ -radiation 4575 independent reflections were measured on an automatic four-circle diffractometer. Using an observed unobserved cutoff at  $2.0\sigma(I)$ , 2250 were recorded as observed. No corrections have been made for absorption on secondary extinction effects.

The structure was solved by direct methods<sup>7</sup> and refined by full-matrix least squares technique.<sup>8,\*</sup> Hydrogen atom positions were calculated and the positional parameters only were refined. Anisotropic temperature factors were introduced for O, N, and C-atoms, and weights in least squares were calculated from the standard deviations in intensities,  $\sigma(I)$ , taken as

$$\sigma(I) = [C_T + (0.02C_N)^2]^{1/2}$$

where  $C_T$  is the total number of counts and  $C_N$  net count (peak minus background). The conventional  $R$ -value arrived at was 6.9% (weighted value  $R_w = 4.9\%$ ) for 2250 observed reflections. The form factors were those of Hanson *et al.*<sup>9</sup> except for hydrogen.<sup>10</sup> The final fractional coordinates and thermal vibration parameters are given in Table 1. The expression for anisotropic vibration is:

$$\exp[-(B_{11}h^2 + B_{22}k^2 + B_{33}l^2 + B_{12}hk + B_{13}hl + B_{23}kl)]$$

\* All programs used are included in this reference.

Table 1. Final fractional coordinates and anisotropic thermal vibration parameters with estimated standard deviations (multiplied by  $10^5$  for non-hydrogens and  $10^4$  for hydrogens). The symbols CC, CM, and OW are used for carbonyl carbons, methyl carbons, and water oxygen, respectively. Hmn is bonded to Cm, HMmn to CMm, and HW to OW.

ATOM	X	Y	Z	B	B11	B22	B33	B12	B13	B23
OM	22986(24)	63964(38)	=5195(62)		343(12)	851(40)	1852(74)	-167(37)	1008(52)	-122(92)
O1	29984(19)	30718(35)	25258(48)		319(12)	672(37)	2318(82)	-207(36)	1186(55)	-819(93)
O2	12689(18)	34963(36)	11789(46)		229(10)	1081(43)	1849(72)	-322(35)	893(49)	247(91)
O3	=1295(17)	46993(37)	-33321(48)		148(10)	1284(47)	2098(75)	-284(36)	448(48)	-642(103)
O4	15093(18)	75269(34)	12893(45)		285(11)	889(48)	1196(64)	-249(35)	662(47)	-198(88)
O5	27596(16)	83599(34)	16178(46)		183(10)	948(42)	1873(74)	43(33)	634(47)	-1174(91)
O6	37588(17)	81549(33)	57259(46)		279(11)	888(46)	2167(77)	150(35)	1286(51)	471(93)
O7	44775(18)	50478(34)	64716(45)		315(12)	790(41)	1634(73)	79(36)	978(51)	579(88)
N1	22578(20)	42332(40)	28744(53)		167(12)	857(48)	1142(76)	-197(39)	518(55)	-133(98)
N2	9208(19)	44288(39)	=15982(50)		164(11)	793(45)	1101(74)	-75(36)	508(52)	-283(93)
N3	7772(19)	71577(37)	-14257(50)		179(11)	644(43)	1079(75)	-84(37)	538(52)	-22(96)
N4	1987(20)	95728(38)	10382(52)		136(10)	503(39)	1286(73)	76(34)	497(49)	-134(99)
N5	36968(19)	96688(38)	42223(48)		143(10)	622(41)	1097(74)	-15(36)	434(49)	144(95)
N6	47087(19)	69497(40)	63186(51)		172(11)	736(45)	1072(75)	85(38)	552(51)	237(96)
N7	36037(20)	46955(40)	31569(50)		186(11)	667(44)	1518(87)	52(39)	745(54)	-206(100)
C1	27483(26)	49186(52)	31315(70)		163(14)	612(52)	1225(93)	15(44)	457(62)	-371(114)
C2	15836(28)	46891(59)	=943(64)		192(14)	855(57)	989(87)	-231(47)	514(63)	252(110)
C3	5876(28)	61740(57)	-12237(78)		177(14)	837(60)	1866(110)	-62(48)	834(67)	-73(134)
C4	14815(26)	88871(51)	=4599(65)		162(14)	564(51)	1281(93)	38(43)	535(65)	170(115)
C5	30598(27)	9919(57)	35208(72)		171(14)	715(54)	1108(99)	-45(45)	468(62)	-514(116)
C6	44997(26)	80747(53)	54948(72)		169(14)	640(53)	1518(101)	-92(45)	684(65)	-192(119)
C7	38150(26)	59170(52)	37659(69)		189(14)	593(51)	1388(95)	-82(44)	674(65)	-39(117)
CC1	31331(25)	41336(50)	29368(61)		173(14)	678(54)	989(86)	1(48)	419(61)	-273(118)
CC2	16842(26)	40621(49)	13971(73)		203(15)	446(48)	1409(99)	-100(44)	715(69)	-171(113)
CC3	4058(25)	50236(50)	-21418(65)		151(14)	766(57)	1447(99)	-115(48)	626(66)	-22(127)
CC4	12519(24)	78058(47)	=1273(65)		156(14)	612(53)	1056(97)	138(45)	415(64)	6(121)
CC5	25714(24)	92391(50)	19709(62)		159(14)	682(56)	1139(91)	-7(45)	517(62)	-146(117)
CC6	3949(23)	86311(46)	51366(62)		134(13)	558(50)	1033(91)	-128(43)	435(59)	-443(109)
CC7	43619(25)	59304(51)	56349(66)		177(14)	715(56)	1364(100)	-15(47)	706(69)	13(125)
CM1	24378(34)	35385(69)	42366(79)		307(18)	1171(71)	1254(103)	-113(60)	765(77)	93(142)
CM2	8375(36)	32943(66)	=23938(91)		331(19)	902(64)	1969(120)	-42(59)	1315(84)	-593(147)
CM3	4469(33)	74813(64)	=13152(75)		234(16)	944(62)	1199(98)	-149(52)	525(68)	-134(133)
CM4	17648(36)	185198(68)	14958(99)		269(17)	811(60)	2518(136)	318(54)	1394(85)	-332(148)
CM5	30901(31)	191208(68)	33858(83)		199(15)	1008(65)	1497(111)	-94(51)	577(78)	346(134)
CM6	51693(33)	70594(69)	80829(74)		259(17)	1202(71)	1240(101)	-15(58)	618(72)	-35(141)
CM7	39923(37)	39942(69)	30259(98)		365(20)	1126(69)	1964(122)	247(61)	1390(89)	152(151)
H11	3843(22)	5222(41)	4308(58)	4.0						
H12	2968(21)	5697(44)	2421(56)	4.0						
H21	1575(24)	4683(50)	575(61)	4.0						
H22	1651(22)	5586(44)	-35(57)	4.0						
H31	95(22)	6324(41)	-1639(56)	4.0						
H32	730(21)	5989(41)	-91(58)	4.0						
H41	1668(20)	8606(41)	-929(54)	4.0						
H42	1139(22)	9502(42)	-1228(55)	4.0						
H51	3012(21)	10862(44)	3328(55)	4.0						
H52	3020(22)	9807(42)	4332(57)	4.0						
H61	4380(21)	7917(42)	4348(58)	4.0						
H62	4849(22)	8665(42)	5974(55)	4.0						
H71	3481(22)	6430(42)	3547(56)	4.0						
H72	3976(21)	6334(42)	3254(54)	4.0						
HM11	2128(22)	3115(44)	4087(55)	4.0						
HM12	2655(22)	4814(42)	5235(59)	4.0						
HM13	2783(22)	2986(44)	4693(56)	4.0						
HM21	407(22)	3291(42)	-3592(59)	4.0						
HM22	914(21)	2558(44)	-1690(57)	4.0						
HM23	1135(21)	3166(44)	-2457(56)	4.0						
HM31	685(21)	8012(45)	-3288(54)	4.0						
HM32	4(23)	7615(42)	-3800(56)	4.0						
HM33	449(22)	6802(45)	-3656(57)	4.0						
HM41	2047(22)	11028(44)	2513(58)	4.0						
HM42	1528(24)	10443(48)	1507(65)	4.0						
HM43	1605(23)	11128(45)	880(58)	4.0						
HM51	1354(23)	10243(43)	4152(57)	4.0						
HM52	3741(21)	10973(46)	3929(56)	4.0						
HM53	3666(22)	9728(44)	2365(59)	4.0						
HM61	5299(22)	6332(44)	8517(56)	4.0						
HM62	5008(21)	7380(44)	8530(56)	4.0						
HM63	5523(23)	7586(43)	8459(55)	4.0						
HW1	3768(22)	3305(45)	2371(56)	4.0						
HW2	4398(23)	3817(42)	4095(58)	4.0						
HW3	4069(22)	4371(44)	2448(58)	4.0						
HW1	2048(22)	6725(47)	-1454(58)	4.0						
HW2	2438(23)	7084(44)	-126(62)	4.0						

The principal axes of the thermal vibration ellipsoids for oxygen, nitrogen, and carbon atoms were calculated from the temperature parameters of Table 1. Maximum root mean squares amplitudes range from about 0.22 Å for ring atoms to about 0.33 Å for methyl carbon atoms and the water oxygen. Due to the size of the molecule, no rigid-body

analysis of translational, librational, and screw motion has been carried out. A comparison between observed and calculated structure factors is presented in Table 2.

Interatomic distances, bond angles and dihedral angles are given in Table 3. The standard deviations, given in parentheses, are estimated from the correlation matrix of the last least



Table 2. Continued.

Table with 24 columns of numerical data and chemical notations (e.g., 26 132 137, K= 3, L=10, K= 89, 72, etc.).



Table 3. Interatomic distances, bond angles and dihedral angles with estimated standard deviations.

DISTANCE	(Å)	DISTANCE	(Å)	DISTANCE	(Å)
O1 = CC1	1,223( 6)	O2 = CC2	1,214( 6)	O3 = CC3	1,219( 5)
O4 = CC4	1,225( 5)	O5 = CC5	1,218( 6)	O6 = CC6	1,223( 6)
O7 = CC7	1,224( 6)	CC1 = N7	1,352( 6)	CC2 = N1	1,333( 6)
CC3 = N2	1,335( 6)	CC4 = N3	1,336( 6)	CC5 = N4	1,342( 6)
CC6 = N5	1,357( 6)	CC7 = N6	1,347( 6)	CC1 = C1	1,531( 7)
CC2 = C2	1,575( 8)	CC3 = C3	1,525( 8)	CC4 = C4	1,524( 7)
CC5 = C5	1,538( 7)	CC6 = C6	1,523( 7)	CC7 = C7	1,531( 7)
C1 = N1	1,449( 7)	C2 = N2	1,436( 6)	C3 = N3	1,436( 7)
C4 = N4	1,461( 6)	C5 = N5	1,457( 6)	C6 = N6	1,456( 6)
C7 = N7	1,446( 6)	N1 = CM1	1,436( 7)	N2 = CM2	1,454( 7)
N3 = CM3	1,454( 7)	N4 = CM4	1,438( 7)	N5 = CM5	1,455( 7)
N6 = CM6	1,440( 7)	N7 = CM7	1,444( 8)	OM = O5	2,833( 6)
OM = O4	2,864( 6)				

ANGLE	(°)	ANGLE	(°)
O1 = CC1 = C1	121,2( 5)	O2 = CC2 = C2	120,1( 5)
O3 = CC3 = C3	119,9( 5)	O4 = CC4 = C4	122,8( 5)
O5 = CC5 = C5	129,3( 5)	O6 = CC6 = C6	128,1( 5)
O7 = CC7 = C7	121,1( 5)	O1 = CC1 = N7	122,3( 5)
O2 = CC2 = N1	123,7( 6)	O3 = CC3 = N2	123,0( 5)
O4 = CC4 = N3	122,8( 5)	O5 = CC5 = N4	123,6( 5)
O6 = CC6 = N5	123,0( 5)	O7 = CC7 = N6	121,5( 5)
CM1 = N1 = C1	117,4( 5)	CM2 = N2 = C2	115,3( 5)
CM3 = N3 = C3	115,6( 5)	CM4 = N4 = C4	117,8( 5)
CM5 = N5 = C5	117,1( 5)	CM6 = N6 = C6	116,0( 5)
CM7 = N7 = C7	116,0( 5)	CM1 = N1 = CC2	118,4( 5)
CM2 = N2 = CC3	118,3( 5)	CM3 = N3 = CC4	123,5( 5)
CM4 = N4 = CC5	123,7( 5)	CM5 = N5 = CC6	121,1( 5)
CM6 = N6 = CC7	118,9( 5)	CM7 = N7 = CC1	118,8( 5)
N7 = CC1 = C1	116,4( 5)	N1 = CC2 = C2	116,2( 5)
N2 = CC3 = C3	117,1( 5)	N3 = CC4 = C4	117,1( 5)
N4 = CC5 = C5	116,9( 5)	N5 = CC6 = C6	116,9( 5)
N6 = CC7 = C7	117,4( 5)	C1 = N1 = CC2	123,2( 5)
C2 = N2 = CC3	125,3( 5)	C3 = N3 = CC4	120,4( 5)
C4 = N4 = CC5	118,2( 4)	C5 = N5 = CC6	116,6( 5)
C6 = N6 = CC7	121,7( 4)	C7 = N7 = CC1	124,5( 5)
CC1 = C1 = N1	112,1( 5)	CC2 = C2 = N2	110,1( 4)
CC3 = C3 = N3	114,1( 5)	CC4 = C4 = N4	111,8( 5)
CC5 = C5 = N5	109,4( 4)	CC6 = C6 = N6	111,7( 5)
CC7 = C7 = N7	111,2( 5)	OM = O5 = CC5	143,6( 4)
OM = O4 = CC4	139,2( 4)	O5 = OM = O4	103,8( 2)

DIMEDRAL ANGLE	(°)	DIMEDRAL ANGLE	(°)
CC1 = C1 = N1 = CC2	86,7( 6)	C1 = N1 = CC2 = C2	1,4( 8)
N1 = CC2 = C2 = N2	179,6( 5)	CC2 = C2 = N2 = CC3	-82,0( 7)
C2 = N2 = CC3 = C3	-9,5( 8)	N2 = CC3 = C3 = N3	-68,7( 7)
CC3 = C3 = N3 = CC4	134,4( 5)	C3 = N3 = CC4 = C4	175,9( 5)
N3 = CC4 = C4 = N4	-175,7( 4)	CC4 = C4 = N4 = CC5	-87,4( 6)
C4 = N4 = CC5 = C5	-179,8( 5)	N4 = CC5 = C5 = N5	168,7( 5)
CC5 = N4 = N5 = CC6	63,8( 7)	C5 = N5 = CC6 = C6	-161,2( 4)
N5 = CC6 = C6 = N6	177,1( 4)	CC6 = C6 = N6 = CC7	-72,1( 8)
C6 = N6 = CC7 = C7	-13,2( 7)	N6 = CC7 = C7 = N7	-163,5( 5)
CC7 = C7 = N7 = CC1	-92,4( 6)	C7 = N7 = CC1 = C1	-8,6( 7)
N7 = CC1 = C1 = N1	179,9( 5)		

Table 4.

Distance (Å)	(I)	(II)	(III)	Cycloheptasarcosyl
CC—C	1.527	1.530	1.525	1.534
CC—N	1.344	1.345	1.346	1.343
CC—O	1.228	1.232	1.224	1.221
C—N	1.456	1.453	1.454	1.449
CM—N	1.483	1.487	1.466	1.446

Table 5.

Angle (°)	(I)	(II)	(III)	Cycloheptasarcosyl
(CM—N—CC) <i>cis</i>	118.7	118.6	117.9	118.6
(CM—N—CC) <i>trans</i>	123.9	123.5	123.8	122.8
(C—N—CC) <i>cis</i>	123.8	122.8	123.5	123.7
(C—N—CC) <i>trans</i>	117.2	117.5	116.1	118.4

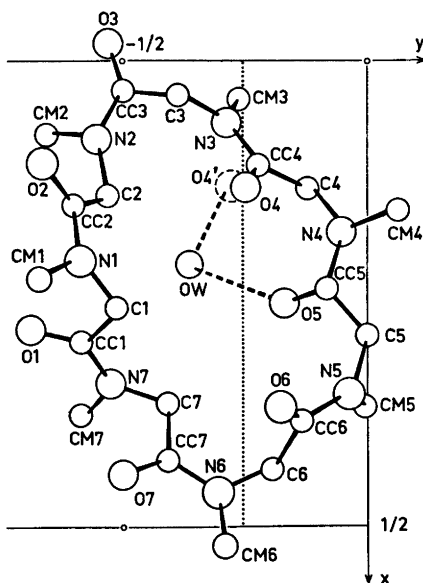


Fig. 1.

squares refinement cycle. Fig 1 shows the molecule viewed along [001].

By averaging bond distances of Table 3, and comparing with the results of the corresponding pentameric <sup>6</sup> (I) and octameric <sup>5</sup> (II) compounds, and with cycloalanyl-tetrasarcosyl <sup>11</sup> (III), no significant differences are observed (Table 4).

The somewhat longer CM—N distances of (I) and (II) are possibly connected with the fact that for these compounds, methyl hydrogens were not included in the calculations. The geometry of the *cis* and *trans* *N*-methyl amide groups, respectively, is also roughly the same (Table 5).

Fig. 1 shows that the ring conformation is *cis,cis,cis,cis,trans,trans,trans*, quite unexpectedly a widely different conformation from that of cyclooctasarcosyl <sup>5</sup> (*cis,cis,trans,trans,cis,cis,trans,trans*).

The water molecules link glide plane equivalent molecules to endless chains along [001]. The two O...O distances are approximately equal ( $OW \cdots O_5 = 2.833 \text{ \AA}$ ;  $OW \cdots O_4' = 2.864 \text{ \AA}$ ) and the angle  $O_5 \cdots OW \cdots O_4'$  is  $103.8^\circ$ .

Since the shortest  $CC \cdots N$  distance across the ring is longer than  $3.5 \text{ \AA}$ , no direct transannular contact can be held responsible for the rigidity of this 21-membered ring. As in earlier

findings,<sup>4-6,11</sup> the explanation must be sought in the intrinsic conformation of the peptide chain itself.<sup>3</sup>

Apart from the hydrogen bonds, there are no short inter-molecular contacts.

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## REFERENCES

1. Dale, J. and Titlestad, K. *Chem. Commun.* (1969) 656.
2. Titlestad, K., Groth, P. and Dale, J. *Chem. Commun.* (1973) 346.
3. Titlestad, K., Groth, P. and Dale, J. *Chem. Commun.* (1973) 646.
4. Groth, P. *Acta Chem. Scand.* 24 (1970) 780.
5. Groth, P. *Acta Chem. Scand.* 27 (1973) 3117.
6. Groth, P. *Acta Chem. Scand.* 27 (1973) 3419.
7. Germain, G., Main, P. and Woolfson, M. M. *Acta Crystallogr. A* 27 (1971) 368.
8. Groth, P. *Acta Chem. Scand.* 27 (1973) 1837.
9. Hanson, H. P., Herman, F., Lea, J. D. and Skillman, S. *Acta Crystallogr.* 17 (1964) 1040.
10. Stewart, R. F., Davidson, E. R. and Simpson, W. T. *J. Chem. Phys.* 43 (1965) 3175.
11. Groth, P. *Acta Chem. Scand. A* 28 (1974) 449.

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