

## Crystal Structure of DL-Tyrosine

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The crystal structure of DL-tyrosine has been determined by X-ray methods using 1180 observed reflections collected by counter methods. The crystals are orthorhombic, space group  $Pna2_1$ , with unit cell dimensions  $a=20.83_6$  Å,  $b=6.81_0$  Å, and  $c=5.90_5$  Å. The structure was solved by direct methods and refined to a conventional  $R$ -factor of 0.037; estimated standard deviations in bond lengths not involving hydrogen are 0.003 Å and in angles 0.2°. The molecular dimensions are found to be nearly identical with those found for L-tyrosine. The difference in crystal packing between the L and the DL form is discussed.

The preferred conformation of tyrosine is of interest in several biochemical relations and has frequently been discussed. Information on the structure of the molecule in solution is difficult to obtain; one approach to the problem is to examine the molecular structure in various crystal modifications in order to study the influence of the environment on the conformation. For an optically active compound crystals of the active and racemic forms may both be examined.

The conformation of tyrosine in the crystals of L-tyrosine<sup>1</sup> is different from that in L-tyrosine hydrochloride<sup>2</sup> but is the same as the one reported for L-phenylalanine hydrochloride<sup>3</sup> and potassium L-tyrosine-*o*-sulphate dihydrate.<sup>4</sup> We have now investigated the crystal structure of DL-tyrosine in which the molecular packing necessarily is different from that in L-tyrosine. Furthermore, as the crystal data on L-tyrosine were found to be different from those derived from powder data,<sup>5</sup> we felt that the crystal data on DL-tyrosine arrived at from powder photographs<sup>6</sup> ought to be checked against single crystal results.

### EXPERIMENTAL

DL-Tyrosine was recrystallized following the same procedure as the one worked out for L-tyrosine.<sup>7</sup> A single crystal with approximate dimensions  $0.2 \times 0.25 \times 0.35$  mm<sup>3</sup> was mounted with the direction of the largest dimension ( $c$ ) along the goniometer head axis and used in all the X-ray experiments.

Precession photographs indicate orthorhombic symmetry; space group extinctions are  $(0kl)$  for  $k+l$  odd and  $(h0l)$  for  $h$  odd. The absence of centers of symmetry was proved by Wilson statistics; the space group is thus  $Pna2_1$ .

Unit cell dimensions were determined from diffractometer measurements on 30 general reflections using  $MoK(\alpha_1 + \alpha_2)$  radiation ( $\lambda = 0.71069 \text{ \AA}$ ). It may be noted that one axis ( $6.810 \text{ \AA}$ ) has half the length of that derived from powder data ( $13.68 \text{ \AA}$ ).<sup>6</sup>

The intensity data were recorded with the use of a SYNTEX P1 diffractometer with graphite crystal monochromated  $MoK$  radiation for 1369 independent reflections with  $2\theta < 60^\circ$  using the  $\omega - 2\theta$  scanning mode with scan speed varying from 2 to  $12^\circ \text{ min}^{-1}$  dependent on the intensity. The scan range was from  $1^\circ$  below  $2\theta(\alpha_1)$  to  $1^\circ$  above  $2\theta(\alpha_2)$  and background counts were taken for half the scan time at each of the scan range limits. Three standard reflections measured after every 50 reflections showed no systematic fluctuation during the experiment. The 1180 reflections greater than  $2.5\sigma(I)$  were considered to be observed; the remaining 189 were excluded from the structure refinement procedure.

The structure determination was based on the program MULTAN written by P. Main, M. M. Woolfson and G. Germain. All other computer programs applied are described in Ref. 10.

Atomic form factors used were those of Doyle and Turner<sup>8</sup> for oxygen, nitrogen, and carbon atoms and of Stewart, Davidson and Simpson<sup>9</sup> for hydrogen.

#### CRYSTAL DATA

DL-4-Hydroxyphenylalanine (DL-tyrosine)  $C_9H_{11}O_3N$ , orthorhombic.  $a = 20.836 (0.0008) \text{ \AA}$ ;  $b = 6.810 (0.002) \text{ \AA}$ ;  $c = 5.905 (0.002) \text{ \AA}$ .  $V = 837.9 \text{ \AA}^3$ ,  $M = 181.19$ ,  $F(000) = 384$ ,  $Z = 4$ ,  $D_{\text{calc}} = 1.435 \text{ g cm}^{-3}$ . Absent reflections:  $(0kl)$  for  $k+l$  odd,  $(h0l)$  for  $h$  odd. Space group  $Pna2_1$ .

#### STRUCTURE DETERMINATION

The intensity data were put on an absolute scale by Wilson's statistical method and normalized structure amplitudes were calculated. The 154 of these with  $E$ -values larger than 1.50 were used as input in the program assembly MULTAN.<sup>11</sup> Of the solutions obtained the one with the highest figure of merit was used as the basis for an  $E$ -map which readily indicated the positions of all non-hydrogen atoms. Initial least-squares refinement included the thirteen heavy atoms with isotropic, then anisotropic temperature factors. When the  $R$ -factor reached 0.05 all hydrogen atoms were assigned coordinates from stereochemical considerations and were included in the further least-squares calculations with individual isotropic thermal parameters. After five cycles the shifts in the parameters were negligible compared to the corresponding standard deviations. The final conventional  $R$ -factor was 0.037 and the weighted  $R$  was 0.055. A comparison of observed and calculated structure factors is given in Table 1; the final parameters for non-hydrogen atoms are listed in Table 2 and for hydrogen atoms in Table 3.

Magnitudes and directions of the principal axes of the ellipsoids of vibrations are given in Table 4. An analysis of the thermal parameters showed that the phenol and alanine parts may each be regarded as rigid bodies. The phenol part has the axis of the largest oscillation ( $4.6^\circ$ ) nearly parallel to the O1-C4-C1-C7 direction; the alanine part has the corresponding axis ( $7.9^\circ$ ) parallel to the C8-O3 direction. The bond lengths were corrected for librational effects according to this model.





Table 2. Fractional atomic coordinates and thermal parameters with estimated standard deviations ( $\times 10^6$ ) for non-hydrogen atoms. The temperature factor is given by  $\exp - (B_{11}h^2 + B_{22}k^2 + B_{33}l^2 + B_{12}hk + B_{13}hl + B_{23}kl)$ .

Atom	<i>x</i>	<i>y</i>	<i>z</i>	<i>B</i> <sub>11</sub>	<i>B</i> <sub>22</sub>	<i>B</i> <sub>33</sub>	<i>B</i> <sub>12</sub>	<i>B</i> <sub>13</sub>	<i>B</i> <sub>23</sub>
O1	-10168	73060	17400	112	2000	1565	-117	-20	273
	7	27		3	39	49	17	22	81
O2	17158	40200	2030	169	1400	1506	-219	-126	-464
	8	24	44	4	31	44	17	22	68
O3	21463	67438	-12859	161	1659	1047	-103	51	30
	8	25	45	4	34	43	18	21	68
N	19866	49323	44313	113	1336	949	36	5	116
	8	26	45	3	35	40	17	21	71
C1	9488	79124	28302	108	933	1319	45	35	-78
	9	27	47	4	34	52	18	26	76
C2	6665	84401	7769	117	1136	1185	39	161	313
	9	30	50	4	37	49	20	26	77
C3	96	82104	4253	122	1182	1284	68	-48	202
	10	31	52	4	36	57	20	28	83
C4	-3696	74333	21334	104	972	1340	-15	31	-252
	9	29	48	4	35	55	19	26	79
C5	-970	68510	41867	122	1119	1247	-13	116	314
	10	31	52	4	35	53	20	26	78
C6	5619	71018	45145	123	1189	1151	117	53	284
	10	30	48	4	37	50	20	27	76
C7	16567	82769	32339	108	1115	1426	-30	18	-331
	9	29	48	4	35	61	19	24	76
C8	20904	65017	27242	87	1181	850	-76	-9	-81
	9	31	47	4	37	45	18	22	74
C9	19741	56741	3502	78	1294	931	54	-58	-460
	8	30	47	3	39	46	18	21	77

Standard deviations in bond lengths and angles were computed from the correlation matrix ignoring standard deviations in cell parameters. For distances and angles the standard deviations were calculated to be 0.003 Å and 0.2°, respectively, except for the bond lengths involving hydrogen where the standard deviation varied from 0.03 to 0.06 Å.

## DISCUSSION

A drawing of the molecule is shown in Fig. 1 in which the numbering of the atoms is also indicated. The bond lengths (corrected for thermal effects) and valence angles are given in this figure; interatomic distances and bond angles are also listed in Table 5.

The unit cell dimensions for DL-tyrosine are quite close to those found for L-tyrosine,<sup>1</sup> and it is interesting to note that the same relation exists between the space groups of L- and DL-tyrosine as that reported for L- and DL-alanine,<sup>12</sup> *viz.*  $P2_12_12_1$  for the optically active crystals and  $Pna2_1$  for the racemate. It has recently been pointed out by Pedone and Benedetti<sup>13</sup> that such relations

Table 3. Fractional coordinates ( $\times 10^4$ ) and isotropic thermal parameters ( $\text{Å}^2$ ) with estimated standard deviations for hydrogen atoms.

Atom	$x$	$y$	$z$	$B$
H2	984	8954	-479	5.3
	23	69	112	1.2
H3	-202	8587	-1020	2.5
	13	41	75	0.6
H5	-367	6261	5327	2.1
	15	40	72	0.6
H6	721	6847	5787	4.5
	15	53	87	0.8
H71	1744	8652	4920	2.4
	17	49	91	0.9
H72	1773	9259	2367	1.6
	12	41	65	0.6
H8	2511	6857	3006	1.4
	15	41	60	0.6
HO1	-1170	6871	2893	5.1
	18	45	85	0.8
HN1	1668	4558	4428	5.8
	22	65	112	1.2
HN2	2018	5535	5782	4.8
	14	45	65	0.6
HN3	2284	3906	4282	3.4
	13	43	71	0.7

Table 4. R.m.s. amplitudes of vibration  $(\overline{u^2})^{\frac{1}{2}}$  and  $B$ -values ( $\text{Å}^2$ ) along the principal axes of vibration given by the components of a unit vector  $\mathbf{e}$  in fractional coordinates ( $\times 10^3$ ).

Atom	$(\overline{u^2})^{\frac{1}{2}}$	$B$	$e_x$	$e_y$	$e_z$
O1	.219	3.80	-9	143	23
	.165	2.15	1	21	-168
	.155	1.89	47	26	6
O2	.208	3.41	-38	89	3
	.181	2.60	20	85	-118
	.143	1.63	21	81	121
O3	.203	3.26	26	-124	4
	.182	2.63	40	79	17
	.135	1.45	5	5	-168
N	.178	2.51	9	144	13
	.157	1.95	-47	28	1
	.129	1.32	0	12	-169
C1	.158	1.97	41	52	66
	.154	1.86	8	79	-140
	.143	1.61	-24	112	68
C2	.176	2.45	31	86	85
	.158	1.98	29	-115	23
	.132	1.37	23	32	-145

Table 4. Continued.

	.173	2.36	29	117	18
C3	.163	2.09	-34	68	91
	.145	1.66	18	-58	142
	.162	2.07	14	-89	124
C4	.150	1.77	45	41	-26
	.144	1.63	-5	109	112
	.171	2.32	33	69	93
C5	.164	2.13	-29	117	13
	.137	1.49	19	56	-141
	.181	2.58	31	107	44
C6	.153	1.85	-37	90	28
	.139	1.53	2	44	-161
	.171	2.32	10	-110	107
C7	.153	1.85	46	1	-52
	.149	1.75	11	98	121
	.169	2.25	13	-141	9
C8	.135	1.45	46	39	-24
	.122	1.18	6	13	167
	.179	2.54	8	138	-51
C9	.132	1.38	-43	39	63
	.120	1.14	21	30	149

may indeed be expected between crystals of the optically active and racemic compounds where virtually identical layers of molecules are packed through the operation of different symmetry elements. This appears to be the case in L- and DL-tyrosine, the layer of molecules between the diagonal glide-planes shown in Fig. 3 is nearly the same in both structures. The layers are differently

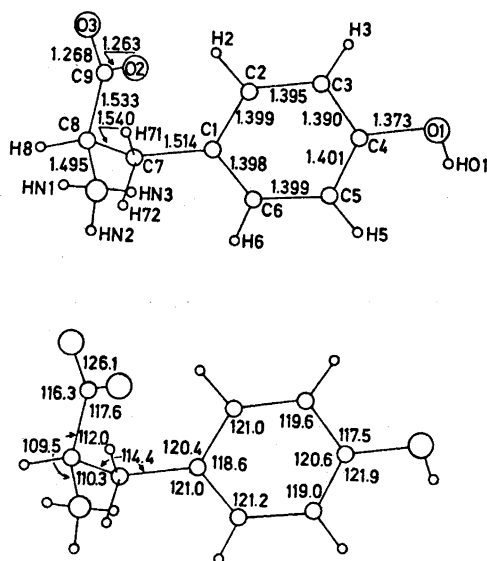


Fig. 1. Bond lengths (Å) corrected for thermal vibration effects and angles (°) in DL-tyrosine.

Table 5. Bond lengths (Å) and bond angles (°) in DL-tyrosine.

	Bond length	Corrected	Bond angle	
C1-C2	1.395	1.399	C1-C2-C3	121.0
C2-C3	1.393	1.395	C2-C3-C4	119.6
C3-C4	1.386	1.390	C3-C4-C5	120.6
C4-C5	1.396	1.401	C4-C5-C6	119.0
C5-C6	1.397	1.399	C1-C6-C5	121.2
C1-C6	1.394	1.398	C2-C1-C6	118.6
C1-C7	1.513	1.514	C7-C1-C2	120.4
C7-C8	1.540		C7-C1-C6	121.0
C8-C9	1.530	1.533	O1-C4-C5	121.9
C8-N	1.485	1.495	O1-C4-C3	117.5
C9-O2	1.251	1.263	C1-C7-C8	114.4
C9-O3	1.262	1.268	C7-C8-C9	112.0
C4-O1	1.371	1.373	C7-C8-N	110.3
O1-HO1	0.81		N-C8-C9	109.5
N-H1N	0.70		C8-C9-O2	117.6
N-H2N	0.90		C8-C9-O3	116.3
N-H3N	0.94		O2-C9-O3	126.1
C2-H2	1.05			
C3-H3	0.99			
C5-H5	0.97			
C6-H6	0.84			
C7-H71	1.04			
C7-H72	0.87			
C8-H8	0.92			
			Hydrogen bond lengths	
			O1-O2( -x, 1-y, $\frac{1}{2}+z$ )	2.668
			N-O1 ( -x, 1-y, $\frac{1}{2}+z$ )	2.875
			N-O3 ( x, y, 1+z)	2.833
			N-O3 ( $\frac{1}{2}-x$ , $-\frac{1}{2}+y$ , $\frac{1}{2}+z$ )	2.856

packed, however, as may be expected since all layers in L-tyrosine are identical whereas the layers in DL-tyrosine consist alternately of D and L forms. The difference in packing is visualized in Figs. 2 (A) and (B) which show the coupling of molecules of different layers through the alanine parts; (A) shows the conditions in L-tyrosine as viewed along the *a*-axis (6.913 Å) and (B) represents the DL-tyrosine structure as viewed down the *b*-axis (6.810 Å). The pertinent symmetry elements responsible for the coupling of layers are screw axes and glide-planes, respectively. The molecular packing is somewhat denser in the racemate, the density being 1.435 g cm<sup>-3</sup> as compared to 1.414 g cm<sup>-3</sup> for L-tyrosine.

The bond lengths and angles found in the present investigation are nearly identical to those determined from the study of the L-form. Except for the C9-O2 bond in the carboxyl group the differences in corresponding bonds are insignificant, 0.008 Å or less; the C9-O2 bond is found 0.017 Å longer than in L-tyrosine. The differences in corresponding angles are 0.7° or less, the largest dissimilarities are found in the alanine moiety. The conformation in the two structures are also equal; the angle between the phenyl ring plane and the plane defined by the C1-C7-C8 atoms is 85.8° (as compared to 85.9° in L-tyrosine), the dihedral angles C1-C7-C8-C9 and C1-C7-C8-N are 52.5° and 290.3, respectively (53.1° and 290.7° in L-tyrosine) whereas the torsional angle N-C8-C9-O2 is 11.7°, 2.3° less than the corresponding angle in L-tyrosine.



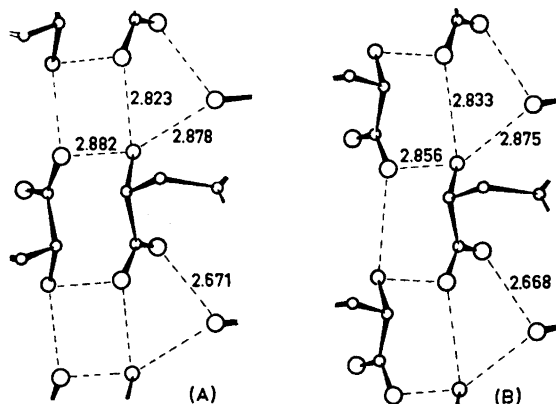


Fig. 2. Molecular arrangement in L-tyrosine (A) and in DL-tyrosine (B) crystals.

The hydrogen bond system in the crystals of DL-tyrosin is shown in Fig. 3. Each molecule is hydrogen donor as well as acceptor in four hydrogen bonds in such a way that one molecule is hydrogen bonded to six neighbouring molecules. The hydrogen bond lengths (given in Table 5) within the layers of molecules of equal chirality are nearly the same as the corresponding bonds in L-tyrosine; the hydrogen bonds linking the layers together are by 0.026 Å shorter in the racemate than in the crystals of the L-form.

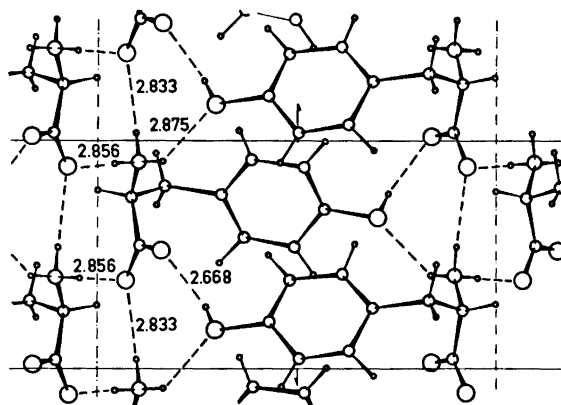


Fig. 3. The crystal structure of DL-tyrosine as viewed along the *b*-axis.

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