

A Refinement of the Crystal Structure of Tellurium Dibzenethiosulphonate

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The crystal structure of tellurium dibzenethiosulphonate, $\text{Te}(\text{S}_2\text{O}_2\text{C}_6\text{H}_5)_2$, has been refined by full-matrix least squares methods, using three-dimensional diffractometer data, to a conventional R value of 0.038. The crystals are orthorhombic, space group $Pbcn$ (No. 60), with $a = 14.467 \text{ \AA}$, $b = 11.137 \text{ \AA}$, $c = 10.455 \text{ \AA}$, and four molecules in the unit cell.

The tellurium atom of the nonplanar $\text{S}-\text{S}-\text{Te}-\text{S}-\text{S}$ chain lies on a two-fold axis, and the molecule thus occurs in a *trans* form. The following dimensions of the chain, with standard deviations in parentheses, have been found: $\text{Te}-\text{S} = 2.380(2) \text{ \AA}$, $\text{S}-\text{S} = 2.080(2) \text{ \AA}$, $\angle \text{S}-\text{Te}-\text{S} = 97.61(6)^\circ$, $\angle \text{S}-\text{S}-\text{Te} = 103.46(7)^\circ$. The dihedral angle between the $\text{S}-\text{S}-\text{Te}$ and $\text{S}-\text{Te}-\text{S}$ planes is 97.4° .

The crystal structure of a *trans* square-planar complex of tellurium dibzenethiosulphonate with trimethylenethiourea¹ was recently reported, and the structures of two similar complexes, one with ethylenethiourea and another with tetramethylthiourea, are under study in this Laboratory. In connection with investigations of this type of complex, it was thought of interest to refine the structure of uncomplexed tellurium dibzenethiosulphonate, which was reported in 1956 by Øyum and Foss.² With more modern facilities, it should be possible to determine the structural parameters more accurately. The lengths of the Te-S and S-S bonds are of special interest for comparison with the corresponding bond lengths in the complexes mentioned above.

EXPERIMENTAL

The yellow orthorhombic crystals of $\text{Te}(\text{S}_2\text{O}_2\text{C}_6\text{H}_5)_2$ were prepared by Foss.³

The data collection was made by means of a Siemens automatic, off-line, single crystal diffractometer (AED), using (Ni-filtered) $\text{CuK}\alpha$ radiation. The diffractometer was operated as a three-circle instrument.

A bipyramidal crystal with the following dimensions, given as distances from a common origin to the faces, was used: distances to $(\bar{1}00)$ and $(100) = 0.046 \text{ mm}$; to $(\bar{1}\bar{1}1)$ and $(11\bar{1}) = 0.061 \text{ mm}$; to $(\bar{1}1\bar{1})$ and $(\bar{1}\bar{1}1) = 0.062 \text{ mm}$; to $(\bar{1}11)$ and $(11\bar{1}) = 0.056 \text{ mm}$; and to (111) and $(111) = 0.059 \text{ mm}$.

The crystal was mounted with the α axis approximately along the ϕ axis of the diffractometer. Its orientation and cell dimensions were determined by measuring θ , χ , and ϕ for three non-coplanar reciprocal vectors, and the setting angles for all reflections could be calculated.

For determination of unit cell dimensions by least squares, θ angles of 8 reflections at high θ were measured.

Intensity data were collected using a scintillation counter and $\theta - 2\theta$ scan technique. The scan-speed was chosen to be 0.5 degrees per minute, with automatic choice of greater scan-speed for strong reflections. When an attenuation filter was needed to avoid counting losses, the proper one was automatically inserted into the primary beam.

Each reflection was scanned between $\theta_1 = \theta - 0.40^\circ$, and $\theta_2 = \theta + 0.40^\circ + 0.15^\circ \text{tg}\theta$, where θ is the Bragg angle for the α_1 peak. The scan was performed by going from θ to θ_1 , then from θ_1 to θ_2 , and finally from θ_2 to θ . The intensities over each scanning range were recorded at the end of each step. The background was measured for one half of the total scan time at θ_1 , and one half at θ_2 . This scan procedure provides a check of the setting angles.

Two reference reflections were measured at intervals of 50 reflections, and the setting angles of these two reflections were checked by horizontal and vertical half shutters.

The net intensities were brought to a common scale by using the intensity variations of the reference reflections. The lower limit for observed reflections was set equal to two times the standard deviation in net intensity. The standard deviation is $(I_t + I_b)^{\frac{1}{2}}$, where I_t is the total intensity, and I_b is the background intensity. Out of 1631 reflections with $\theta < 70^\circ$, 1373 were found stronger than the lower limit. The remaining 258 reflections were assigned an intensity equal to the limit, and labelled as unobserved reflections.

Absorption ($\mu = 193 \text{ cm}^{-1}$), Lorentz, and polarization corrections were carried out. A modified version of the absorption correction method described by Busing and Levy⁴ was applied, using a $6 \times 8 \times 10$ grid.

The calculated structure factors were based on the scattering curve of Stewart *et al.*⁵ for hydrogen, and the curves listed in *International Tables*⁶ for all other atoms. Using the $\Delta f'$ and $\Delta f''$ values given by Cromer,⁷ the tellurium and sulphur scattering curves were corrected for anomalous dispersion by taking the amplitude of f as the corrected value.

The structure was refined by a least squares, full-matrix program minimizing the function

$$r = \sum W(|F_o| - |F_c|)^2$$

Table 1. Computer programs.

Name	Function	Origin
CELL	Cell dimensions	Weizmann Institute of Science, Rehovoth, Israel
BDLS	Least squares refinement	
DIAN	Distances and angles	
ABCD	Absorption correction	As above. Modified by the present author, to take into account a general orientation of the crystal
AED1	Setting angles for the diffractometer	Mr. Knut Maartman-Moe, Chemical Institute, University of Bergen, Norway
PLAN	Weighted least squares planes	
DAT1	Data processing	Present author
ASEN	Fourier summation	
ZACH	Extinction correction	

where K is a scale factor, and the weight, W , is the inverse of the variance of F_o . The variance of F_o is

$$\sigma^2(F_o) = F_o^4[I_t + I_b + k^2(I_t - I_b)^2]/4(I_t - I_b)^2$$

where k may be interpreted as the relative standard deviation in the scaling curve based on reference reflections. It was given the value 0.008. Non-observed reflections for which $K|F_c|$ is greater than the observable limit, are included in the refinement with $|F_o|$ equal to the observable limit.

The calculations were carried out on an IBM 360/50 H computer. The programs used are listed in Table 1.

CRYSTAL DATA

The crystals of tellurium dibenzenethiosulphonate are orthorhombic. Based on 8 high-angle θ -values measured on the diffractometer, a least squares procedure gave the following cell dimensions, with standard deviations in parentheses:

$$a = 14.467(1) \text{ \AA}; b = 11.137(1) \text{ \AA}; c = 10.455(1) \text{ \AA}; V = 1684.5(3) \text{ \AA}^3.$$

Other pertinent values from the earlier study of Øyum and Foss² are

$$M = 474.07; F(000) = 1020; Z = 4.$$

Systematic absences are $h\bar{k}0$ when $h+k$ is odd, $h0l$ when l is odd, $0kl$ when k is odd. The uniquely determined space group is $Pbcn$ (No. 60).

REFINEMENT OF THE STRUCTURE

With four molecules in the unit cell, a two-fold rotation axis must pass through the tellurium atom. Using the positional and thermal parameters given by Øyum and Foss² for the non-hydrogen atoms, refinement by least squares was started. Refinement on the scale factor, positional parameters and individual thermal parameters resulted in an R value of 0.065. Introduction of anisotropic thermal parameters for the tellurium and sulphur atoms lowered the R value to 0.051, and introduction of anisotropic thermal parameters for the rest of the atoms lowered it to 0.040. Applying the test of Hamilton,⁸ both changes in the model were found to be significant at the 0.5 % level.

There seemed to be a marked tendency of $|F_o|$ values for strong, low-order reflections to be smaller than the corresponding $K|F_c|$ values. A correction for extinction was therefore carried out, using the expression given by Zachariasen.⁹ The absorption term in this expression was set equal to 1. With I_o on an absolute scale, the value of the extinction parameter, C , was found to be 4.7×10^{-7} . Additional refinement on the parameters mentioned above lowered the R value to 0.038.

A three-dimensional difference Fourier summation at this point showed peaks between 0.40 and 0.67 e/ \AA^3 at the positions where the hydrogen atoms would be expected. The hydrogen positions were taken from the difference map, and a refinement with hydrogen included was attempted. This attempt was not successful, however, as the hydrogen atoms were shifted to

unreasonable positions, with C—H bond lengths between 0.81 and 1.12 Å, and one of the C—C—H bond angles as low as 99°. The hydrogen atoms were therefore left out of the structure factor calculations and the refinement.

In a final full-matrix, least squares cycle no shift was greater than 0.4 times the standard deviation. The final *R* value, with non-observed reflections included when $K|F_c|$ exceeds the observable limit, is 0.038. The final atomic coordinates are listed in Table 2, and the thermal parameters in Table 3. The structure factors based on the final parameters are listed in Table 4.

Table 2. Atomic coordinates in fractions of cell edges. Standard deviations from the least squares refinement in parentheses.

	<i>x</i>	<i>y</i>	<i>z</i>
Te	0.0	0.04397 (5)	0.25
S(2)	-0.06880 (10)	0.18473 (14)	0.39244 (16)
S(1)	0.03909 (11)	0.23095 (13)	0.51441 (15)
O(1)	-0.0035 (4)	0.3132 (4)	0.6018 (5)
O(2)	0.0832 (4)	0.1246 (4)	0.5619 (5)
C(1)	0.1187 (4)	0.3095 (5)	0.4153 (6)
C(2)	0.2036 (5)	0.2556 (6)	0.3895 (8)
C(3)	0.2623 (5)	0.3167 (7)	0.3050 (9)
C(4)	0.2353 (5)	0.4259 (7)	0.2530 (10)
C(5)	0.1503 (5)	0.4778 (6)	0.2851 (8)
C(6)	0.0909 (5)	0.4203 (6)	0.3679 (7)

Table 3. Anisotropic thermal parameters (\AA^2) in the form $\exp[-2\pi^2(h^2a^{-2}U_{11} + \dots + 2hka^{-1}b^{-1}U_{12} + \dots)]$. All values have been multiplied by 10^4 . Standard deviations in parentheses.

	U_{11}	U_{22}	U_{33}	U_{12}	U_{23}	U_{13}
Te	426 (3)	396 (3)	402 (3)	0	0	4 (3)
S(2)	391 (7)	453 (8)	442 (9)	38 (6)	-14 (7)	68 (7)
S(1)	615 (8)	428 (7)	343 (8)	-25 (7)	10 (7)	10 (7)
O(1)	1063 (35)	625 (26)	469 (27)	-39 (33)	-151 (22)	189 (33)
O(2)	751 (29)	551 (25)	574 (30)	-19 (23)	182 (24)	-129 (25)
C(1)	493 (31)	412 (28)	436 (33)	-65 (25)	-20 (27)	-32 (27)
C(2)	514 (34)	554 (37)	670 (47)	38 (30)	-31 (37)	-56 (34)
C(3)	490 (37)	700 (49)	909 (58)	-57 (34)	82 (42)	34 (37)
C(4)	596 (37)	746 (46)	900 (54)	-199 (32)	-21 (59)	49 (50)
C(5)	641 (38)	509 (41)	789 (64)	-115 (30)	28 (34)	66 (36)
C(6)	669 (38)	387 (29)	589 (44)	-39 (27)	30 (31)	-9 (33)

RESULTS

The tellurium dibzenethiosulphonate molecule has been described by Öyüm and Foss.² Here the new results will be recorded, and the differences, although not great, between the present investigation and the older one, will be pointed out.

Bond lengths and angles, as calculated from the coordinates of Table 2, are listed in Table 5. Coordinate covariances and standard deviations in unit cell dimensions have been neglected. A drawing of the molecule, with the most important bond lengths and angles, is presented in Fig. 1.

Table 4. Observed and calculated structure factors ($\times 10$). Unobserved reflections are indicated by a minus sign on $F(O)$ and included at the threshold values.

H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)
2	0	C	3432	3114	4	12	0	348	-343	6	6	1	758	-793	7	0	2	313	-311
4	0	C	2240	2246	6	12	0	401	-399	7	6	1	592	-590	8	0	2	1590	-1593
6	0	C	345	347	8	12	0	326	-315	8	6	1	585	-577	6	0	2	366	354
8	0	C	751	756	1	12	0	256	-271	9	6	1	325	-329	10	0	2	386	344
10	0	C	598	579	3	13	0	205	-206	10	6	1	646	-648	11	0	2	512	507
12	0	C	748	740	5	13	0	155	-144	11	6	1	196	-199	12	0	2	974	-943
14	0	C	628	651	1	1	1	905	-954	12	6	1	593	-576	13	0	2	186	176
16	0	C	444	451	2	1	1	788	802	13	6	1	57	-64	14	0	2	611	-596
1	1	O	1752	1552	3	1	1	69	25	14	6	1	461	-453	15	0	2	52	-51
3	1	O	2143	2253	4	1	1	202	227	15	6	1	67	-52	16	0	2	398	-381
5	1	O	1117	1160	5	1	1	94	46	1	7	1	1199	-1229	17	0	2	133	-130
7	1	O	567	606	6	1	1	307	299	2	7	1	477	-472	1	1	2	181	185
9	1	O	521	516	7	1	1	173	127	3	7	1	848	-860	4	0	2	1077	-1061
11	1	O	784	753	8	1	1	279	303	4	7	1	613	-606	3	1	2	1863	-1848
13	1	O	607	620	9	1	1	99	-83	5	7	1	506	-492	4	1	2	705	-676
15	1	C	525	521	10	1	1	504	489	6	7	1	536	-554	5	1	2	1521	-1510
17	1	O	323	327	11	1	1	225	-228	7	7	1	560	-568	6	1	2	52	-51
C	2	C	785	-788	12	1	1	354	343	8	7	1	57	-8	7	1	2	1257	-1301
2	2	O	252	E6	13	1	1	240	-223	9	7	1	355	-350	8	1	2	197	194
4	2	C	426	A0	14	1	1	59	53	10	7	1	145	-145	1	1	2	1027	1026
6	2	O	133	129	15	1	1	210	-222	11	7	1	456	-459	10	0	2	237	235
8	2	O	1346	1367	16	1	1	128	-132	12	7	1	122	-121	11	0	2	799	-788
10	2	C	885	611	17	1	1	65	-104	13	7	1	444	-443	12	1	2	57	87
12	2	O	756	820	0	2	1	379	394	14	7	1	71	-15	13	1	2	467	-452
14	2	C	432	424	1	2	1	311	-346	15	7	1	353	-356	14	1	2	64	-17
16	2	O	205	2C2	2	2	1	1654	-1793	0	8	1	768	-790	15	1	2	429	-419
1	3	C	124	134	3	2	1	191	-161	1	8	1	61	-69	16	1	2	76	-71
3	3	O	118	112	4	2	1	781	-854	2	8	1	725	-736	17	1	2	303	-304
5	3	O	119	120	5	2	1	149	-149	3	9	1	321	-317	9	2	2	317	-314
7	3	C	1455	121	6	2	1	966	-1002	4	8	1	644	-687	1	9	2	114	-71
9	3	O	128	216	7	2	1	191	192	5	8	1	270	-278	2	2	2	2182	2276
11	3	C	646	650	8	2	1	236	-245	6	8	1	422	-425	12	2	2	329	-306
13	3	O	120	120	9	2	1	140	140	7	7	1	258	-251	4	2	2	1489	-1504
15	3	O	190	172	10	2	1	452	-466	8	9	1	456	-458	5	2	2	459	440
17	3	C	172	168	11	2	1	190	183	9	8	1	161	-147	6	2	2	1156	-1226
0	4	O	1265	1267	12	2	1	337	-350	10	8	1	453	-440	7	2	2	187	159
4	4	O	743	742	13	2	1	187	113	11	9	1	651	-664	12	2	2	755	757
6	4	O	1008	1242	14	2	1	426	-417	12	8	1	321	-317	9	2	2	156	-157
8	4	O	1109	1114	15	2	1	50	4	13	8	1	44	-11	10	2	2	496	-398
10	4	C	297	323	16	2	1	266	-264	14	8	1	229	-226	11	2	2	366	-373
12	4	O	102	54	17	2	1	174	-92	1	9	1	266	-264	12	2	2	329	-306
14	4	C	143	-115	1	3	1	1268	-1335	2	9	1	231	-224	13	2	2	65	-66
16	4	C	115	100	2	3	1	50	20	3	9	1	346	-354	14	2	2	329	-315
18	4	O	72	72	3	3	1	1159	-1190	4	9	1	290	-279	15	2	2	49	-56
1	5	C	606	617	4	3	1	151	-159	5	9	1	652	-664	16	2	2	241	-242
3	5	C	573	622	5	3	1	1092	-1114	6	9	1	361	-352	17	2	2	117	116
5	5	C	74	57	6	3	1	543	-549	7	9	1	565	-560	1	3	2	1291	-1341
7	5	C	600	-572	7	3	1	952	-993	8	9	1	87	-61	9	2	2	855	909
9	5	O	262	-262	8	3	1	517	-522	9	9	1	361	-356	3	3	2	1459	-1465
11	5	C	61	-65	9	3	1	776	-800	10	9	1	93	-98	1	2	2	1007	973
13	5	O	212	203	10	3	1	417	-431	11	9	1	219	-207	5	3	2	676	-682
15	5	C	91	52	11	3	1	480	-470	12	9	1	98	-109	6	3	2	199	230
17	5	O	59	-54	12	3	1	151	-159	13	9	1	152	-145	14	3	2	372	-369
2	6	O	80	-80	14	3	1	131	-140	0	10	1	173	-168	8	2	2	211	-212
4	6	O	358	-371	14	3	1	135	127	1	10	1	205	-222	9	3	2	346	-350
6	6	O	226	-159	15	3	1	155	-154	2	10	1	77	-44	10	3	2	345	-327
8	6	C	81	-88	16	3	1	91	93	3	10	1	476	-480	11	3	2	471	-491
10	6	C	97	-88	17	3	1	205	-217	4	10	1	272	-268	12	3	2	173	-175
12	6	O	213	217	0	4	1	469	-492	5	10	1	395	-395	13	3	2	301	-290
14	6	O	209	213	1	4	1	295	313	6	10	1	150	-156	14	3	2	88	-99
16	6	O	504	-525	2	4	1	708	-772	7	10	1	187	-197	15	3	2	265	-295
18	6	O	52	-51	3	4	1	249	241	8	10	1	360	-360	16	3	2	174	-175
7	7	C	481	-480	4	4	1	1045	-1119	5	11	1	300	-300	0	10	1	177	-176
7	7	C	293	-286	5	4	1	449	-449	10	10	1	183	-176	1	3	2	173	-170
9	7	C	143	120	6	4	1	1207	-1216	11	10	1	43	-60	1	4	2	343	-339
11	7	C	54	-39	7	4	1	282	-275	12	10	1	89	-88	2	4	2	227	-257
13	7	C	48	-40	8	4	1	1097	-1113	11	11	1	61	-18	2	4	2	227	-227
15	7	C	112	-162	9	4	1	285	-279	21	11	1	61	-18	4	4	2	120	8
17	7	C	879	10	1	741	719	-719	3	11	1	80	-80	10	10	1	380	370	
19	7	C	604	614	12	4	1	346	-337	5	11	1	57	-57	12	4	2	533	-523
21	7	C	476	-461	12	4	1	346	-337	6	11	1	57	-57	12	4	2	533	-523
8	8	O	100	-99	13	4	1	168	163	6	11	1	114	-122	8	4	2	485	-490
8	8	O	376	-376	14	4	1	218	-202	7	11	1	108	-114	9	4	2	183	-187
10	8	O	283	-277	15	4	1	58	-57	8	11	1	48	-24	10	4	2	544	-538
12	8	C	433	-432	16	4	1	226	-240	9	11	1	53	-36	11	4	2	556	-548
14	8	C	320	-331	1	5	1	1535	-1422	10	11	1	40	-2	12	4	2	488	-493
1	9	C	321	-322	2	5	1	397	412	12	12	1	13	-37	13	4	2	75	-54
3	9	C	317	-329	3	5	1	374	-374	12	12	1	13	-37	13	4	2	874	-875
5	9	C	389	4	5	1	404	-395	3	13	1	176	-173	0	5	2	86	-81	
7	9	C	378	-386	5	5	1	1277	-1313	3	12	1	283	-278	16	4	2	164	-173
9	9	C	382	-377	6	5	1	631	611	4	12	1	52	-14	1	5	2	74	-80
11	9	C	401	-356	7	5	1	924	-950	5	12	1	236	-235	2	5	2	407	-414
13	9	C	391	-387	8	5	1	266	247	6	12	1	163	-147	3	5	2	63	-612
10	10	C	449	-412	9	5	1	931	-935	7	12	1	131	-125	4	5	2	3	

Table 4. Continued.

H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)
6	1	3	688	665	13	7	3	274	257	13	2	4	-55	-40	12	9	4	-39	-39
7	1	3	886	864	14	7	3	99	8	14	2	4	692	683	9	10	4	729	-729
8	1	3	307	256	10	8	3	406	403	15	2	4	51	-37	1	10	4	-62	-22
9	1	3	406	45	1	8	3	84	-72	16	2	4	417	425	2	10	4	711	-722
10	1	3	-55	65	1	8	3	544	510	1	3	4	822	825	3	10	4	-59	19
11	1	3	27	224	3	8	3	235	-28	2	3	4	727	721	4	10	4	400	-477
12	1	3	123	-115	4	8	3	644	653	3	3	4	811	797	6	10	4	343	-355
13	1	3	57	51	5	8	3	292	-281	6	3	4	373	393	6	10	4	343	-355
14	1	3	-52	52	5	8	3	521	512	5	3	4	280	278	7	10	4	-57	-57
15	1	3	121	-116	7	8	3	192	-180	6	3	4	103	112	8	10	4	297	-293
16	1	3	55	-42	8	8	3	499	499	7	3	4	130	91	9	10	4	60	-51
17	1	3	-55	-12	9	8	3	96	-85	8	3	4	85	71	10	10	4	233	-233
0	2	3	395	400	10	8	3	382	388	9	3	4	251	237	1	11	4	700	-702
2	2	3	605	555	11	8	3	97	-88	10	3	4	545	526	2	11	4	-56	-26
2	2	3	782	732	12	8	3	307	311	12	3	4	532	529	11	11	4	587	-580
3	2	3	546	520	13	8	3	-44	1	12	3	4	141	-141	4	11	4	-54	7
4	2	3	157	52	1	9	3	786	767	13	3	4	528	528	5	11	4	389	-383
5	2	3	495	481	2	9	3	-62	-11	14	3	4	-50	-1	6	11	4	-49	25
6	2	3	605	608	3	9	3	362	368	15	3	4	414	417	7	11	4	181	-165
7	2	3	77	-68	4	9	3	129	113	16	3	4	-69	2	8	11	4	-43	49
8	2	3	605	557	5	5	3	-61	-22	0	4	4	1051	1050	0	12	4	417	-419
9	2	3	87	-86	6	9	3	-59	-1	1	4	4	1050	1050	1	12	4	-51	1
10	2	3	61	31	7	9	3	142	128	2	4	4	1115	1155	2	12	4	-52	-352
11	2	3	124	-137	8	9	3	96	84	3	4	4	82	86	3	12	4	-48	4
12	2	3	173	178	9	9	3	129	113	4	4	4	550	558	4	12	4	282	-278
13	2	3	-55	11	10	9	3	96	89	5	4	4	63	47	5	12	4	-44	-5
14	2	3	70	74	11	9	3	210	188	6	4	4	445	444	6	12	4	255	-251
15	2	3	-48	27	12	9	3	93	103	7	4	4	257	-230	1	5	5	238	214
16	2	3	60	57	0	10	3	555	542	8	4	4	255	262	2	1	5	340	340
17	2	3	32	32	3	9	3	-62	1	9	4	254	-216	3	1	5	197	-197	
1	3	1190	1258	2	10	3	400	400	10	4	4	86	-71	9	10	4	990	9	
2	3	473	-470	3	10	3	78	80	11	4	4	-57	-42	5	1	5	449	-410	
3	3	457	534	4	10	3	359	345	12	4	4	151	130	6	1	5	657	637	
4	3	681	-674	5	10	3	85	79	13	4	4	-52	23	7	1	5	550	-546	
5	3	778	764	6	10	3	63	-6	14	4	4	84	77	8	1	5	421	406	
6	3	444	-424	7	10	3	231	223	15	4	4	-67	27	9	1	5	396	-372	
7	3	391	413	8	10	3	-31	-51	16	4	4	54	55	10	1	5	180	-179	
8	3	331	-37	9	10	3	167	163	1	5	4	199	199	11	1	5	157	-157	
9	3	441	441	7	12	3	76	62	4	4	5	159	-153	9	1	5	181	-181	
10	3	122	-120	11	10	3	100	94	3	5	4	428	437	13	1	5	137	-129	
11	3	474	454	1	11	3	-59	66	4	5	4	228	-197	14	1	5	59	-59	
12	3	256	14	2	11	3	-58	28	5	5	4	704	676	15	1	5	86	-87	
13	3	498	493	3	11	3	-56	3	6	5	4	179	-174	16	1	5	49	-27	
14	3	51	41	3	11	3	-56	18	7	5	4	519	512	0	2	5	241	-173	
15	3	362	367	5	11	3	-54	47	8	5	4	-58	-6	1	2	5	150	-144	
16	3	267	172	6	11	3	-51	5	6	4	58	34	2	2	5	517	-233		
1	4	1800	1772	7	11	3	68	45	10	4	4	-58	-69	2	2	5	667	-620	
1	4	3	249	-231	8	11	3	97	99	11	5	4	-67	-55	4	2	5	263	-229
2	4	3	1655	1513	9	11	3	-42	37	12	5	4	65	60	8	2	5	176	-188
3	4	3	156	-171	0	12	3	309	-309	13	5	4	149	-149	6	2	5	731	-727
4	4	3	1278	1282	1	12	3	-54	17	14	5	4	66	63	7	5	5	83	-83
5	4	3	557	-536	2	12	3	234	-227	15	5	4	159	-155	8	2	5	493	-475
6	4	3	662	652	3	12	3	121	-115	6	4	4	704	704	9	2	5	171	-164
7	4	3	207	-187	4	12	3	121	-115	1	6	4	704	-112	10	2	5	201	-202
8	4	3	425	354	5	11	3	100	-95	2	6	4	674	-664	11	2	5	92	-92
9	4	3	75	-83	6	12	3	-54	10	3	6	4	-58	83	12	2	5	273	-220
10	4	3	614	612	7	12	3	122	-130	4	6	4	365	-347	13	2	5	131	-129
11	4	3	122	55	1	13	3	201	-210	5	6	4	-57	-58	14	2	5	175	-174
12	4	3	731	723	2	13	3	-63	-7	6	6	4	101	103	15	2	5	44	-44
13	4	3	88	77	3	13	3	132	-133	7	6	4	64	66	16	1	5	109	-132
14	4	3	586	592	0	0	4	1004	1013	8	6	4	117	98	1	3	5	857	-880
15	4	3	61	29	1	0	4	564	564	9	6	4	64	64	2	2	5	251	-245
16	4	3	354	364	11	0	4	342	328	5	7	4	-332	-309	12	3	5	555	-515
17	4	3	200	211	12	0	4	273	280	6	7	6	69	68	13	3	5	272	-267
18	5	3	395	356	13	0	4	-55	30	7	7	4	295	-287	14	3	5	48	-18
19	5	3	93	82	14	0	4	94	52	8	7	4	135	-125	6	3	5	107	-1065
20	5	3	498	492	15	0	4	-49	-39	9	7	4	163	-166	0	4	5	425	-405
21	5	3	-51	-62	16	0	4	132	113	10	7	4	74	85	1	4	5	252	-264
22	5	3	298	316	1	4	4	1076	1076	11	7	4	74	85	3	0	5	130	-117
23	5	3	526	514	12	0	4	243	-240	12	4	4	100	-97	14	2	5	174	-175
24	5	3	236	227	3	1	4	1558	1514	13	7	4	-64	-3	1	4	5	703	-691
25	6	3	766	759	4	1	4	431	-389	14	7	4	-63	-44	5	5	4	505	-454
26	6	3	242	234	5	1	4	119	1172	0	8	4	160	154	6	5	6	663	-666
27	6	3	888	906	6	1	4	130	-137	1	8	4	-62	-10	7	4	5	221	-218
28	6	3	346	327	7	1	4	1171	1151	2	8	4	187	150	8	5	5	535	-537
29	6	3	1267	1258	8	1	4	166	171	3	8	4	143	-153	9	4	5	99	-95
30	7	2	271	246	9	1	4	974	952	4	8	4	164	-168	10	5	5	578	-571
31	7	2	901	901	10	0	4	73	48	5	8	4	105	-105	11	4	5	176	-176
32	7	2	134	133	11	0	4	749	728	6	8	4	517	-517	12	4	5	422	-417
33	7	2	615	633	12	1	4	-56	17	2	9	4	-61	-16	6	5	5	109	-101
34	7	2	-56	13	1	4	344	339	8	8	4	487	-479	13	4	5	197	194	
35	7	2	277	266	14	1	4	-52	10	9	4	94	-68	15	5	5	324	-323	
36	6	3	59	-53	15	1	4	281	285	10	8	4	362	-358	1	5	5	883	-895
37	6	3	156	184	16	1	4	-43	-3	11	8	4	120	-120	2	5	5	88	-87
38	6	3	-70	-22	0	2	4	164	168	12	8	4	159	-159	10	5	5	521	-523
39	7	2	305	-259	8	2	4	768	745	7	9	4	562	-551</td					

Table 4. Continued.

H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	H	K	L	F(O)	F(C)	
2	2	6	556	-543	0	10	6	248	249	7	7	7	242	241	3	5	8	190	191	
3	2	6	147	-142	1	10	6	65	80	8	7	7	60	-90	4	5	9	210	215	
4	2	6	1115	-1126	2	10	6	256	267	9	7	7	256	256	5	5	8	377	-370	
5	2	6	112	-113	6	10	6	450	452	10	7	7	47	47	6	5	9	172	189	
6	2	6	292	-284	5	10	6	103	106	11	7	7	275	291	7	5	8	154	-317	
7	2	6	624	-625	6	10	6	351	352	1	6	7	-66	0	8	5	8	100	-406	
8	2	6	251	-251	7	10	6	65	53	'2	8	7	281	282	10	5	8	-51	10	
9	2	6	257	-254	8	10	6	288	298	3	8	7	105	-83	11	5	8	111	-352	
10	2	6	214	-223	1	11	6	195	197	4	6	7	334	334	12	5	8	50	1	
11	2	6	254	-246	2	11	6	159	171	5	8	7	156	-160	0	6	8	-84	2	
12	2	6	-5	0	3	11	6	224	225	6	8	7	156	-160	-249	3	6	9	-32	
13	2	6	134	-125	4	11	6	201	201	7	8	7	130	-127	2	6	8	-41	4	
14	2	6	-56	-32	5	11	6	284	306	8	6	7	328	330	3	6	8	195	5	
15	2	6	702	-718	6	11	6	117	107	9	6	7	80	-96	4	6	8	165	51	
16	3	6	233	-253	0	12	6	326	314	10	8	7	293	297	5	6	8	281	444	
17	3	6	737	-752	1	12	6	466	-15	1	9	7	201	203	6	6	8	-78	8	
18	3	6	117	-106	1	7	9	926	882	2	9	7	-61	1	7	9	161	-66		
19	3	6	846	-851	2	11	7	46	46	3	7	7	304	308	6	8	9	105	-318	
20	3	6	88	-89	3	11	7	278	269	4	9	7	22	22	0	6	8	-46	2	
21	3	6	854	-842	4	7	7	234	210	5	9	7	418	432	10	6	8	-50	3	
22	3	6	120	-91	5	1	7	-68	-6	6	9	7	-55	43	11	6	8	67	-206	
23	3	6	437	-436	6	1	7	446	432	7	9	7	354	367	1	7	8	346	-193	
24	3	6	179	-178	7	1	7	339	-332	8	9	7	51	47	2	7	8	-65	7	
25	3	6	290	-282	8	1	7	285	278	9	9	7	233	251	3	7	8	274	-334	
26	3	6	297	-266	9	1	7	101	93	0	10	7	133	125	4	7	8	64	-463	
27	3	6	110	-56	10	1	7	156	160	1	10	7	-57	30	5	7	112	-130		
28	3	6	-51	5	11	1	41	41	2	9	7	-41	5	1	9	157	-37			
29	3	6	79	-77	12	7	58	13	3	10	7	93	93	7	7	8	70	-50		
30	4	6	83	34	12	7	124	124	4	10	7	172	176	10	7	8	-46	3		
31	4	6	-65	6	0	2	1009	954	6	10	7	61	80	4	8	9	204	-221		
32	3	5	93	77	1	2	7	-70	1	7	10	7	114	123	0	8	9	367	-89	
33	4	6	210	-207	2	2	7	889	874	1	11	7	90	26	2	9	8	326	-165	
34	4	6	233	-236	3	2	7	167	-150	2	11	7	92	93	3	9	8	376	-116	
35	4	6	420	-420	4	2	7	353	328	3	11	7	51	22	4	10	8	365	-363	
36	4	6	207	-207	5	2	7	458	459	4	11	7	-55	42	4	10	8	357	-123	
37	4	6	394	-433	6	2	7	235	235	0	8	8	1514	1391	5	8	9	175	-35	
38	4	6	83	7	2	7	79	77	1	0	8	137	-121	6	8	9	87	-57		
39	4	6	94	83	3	2	7	163	-161	10	0	8	349	357	7	8	9	119	-200	
40	4	6	176	161	2	3	7	-61	-61	2	0	8	1254	1183	7	8	9	110	10	
41	5	5	65	-45	3	3	7	500	502	11	0	8	243	-258	7	9	8	111	-241	
42	5	5	97	65	4	3	7	238	-221	12	0	8	314	308	9	8	9	-45	55	
43	7	5	82	72	5	7	490	486	13	0	8	92	-92	10	8	9	-51	13		
44	8	5	130	-106	6	3	7	231	-342	1	1	8	78	749	2	9	8	257	-464	
45	5	5	-43	25	7	7	118	-112	11	1	8	360	359	8	9	8	-50	4		
46	5	5	185	-184	8	3	7	231	-230	3	1	8	790	767	4	10	8	-105	8	
47	5	5	113	-127	9	3	7	559	583	4	1	8	90	42	1	9	117	-105	112	
48	5	5	124	-136	10	3	7	-80	-105	5	1	8	682	683	2	1	9	105	-515	
49	5	5	159	-111	3	3	7	445	452	6	1	8	-70	-2	3	1	9	-35	8	
50	5	5	-55	-12	5	7	-56	7	8	8	614	605	4	1	9	142	-160			
51	0	6	452	-445	13	3	7	327	341	8	1	8	-66	15	5	1	9	159	-31	
52	1	6	322	-322	14	3	7	238	221	16	0	8	91	383	1	9	126	-286		
53	2	6	446	-445	7	1	7	141	151	10	1	8	-42	5	1	9	152	-314		
54	3	6	337	341	1	4	7	118	-112	11	1	8	360	359	2	1	9	220	125	
55	4	6	-67	44	2	4	7	317	327	12	1	8	-54	80	1	9	152	-256	728	
56	4	6	298	306	3	4	7	162	-160	13	1	8	270	282	10	1	9	262	271	
57	6	6	85	55	4	4	7	443	439	0	2	8	252	216	11	1	9	228	-560	
58	7	6	88	78	5	4	7	142	-146	1	2	8	72	72	4	2	9	149	-226	
59	8	6	171	185	6	4	7	839	840	2	2	8	269	257	0	2	9	142	-226	
60	9	6	445	-445	13	3	7	327	341	8	1	8	-66	15	5	1	9	159	-31	
61	9	6	322	-322	14	3	7	238	221	16	0	8	91	383	1	9	126	-286		
62	9	6	446	-445	7	1	7	141	151	10	1	8	-42	5	1	9	152	-314		
63	10	6	155	-125	12	7	7	123	127	4	2	8	43	43	1	2	9	227	-213	
64	11	6	65	-89	9	4	7	-64	-73	5	2	8	95	97	3	9	122	-225		
65	12	6	81	93	10	4	7	536	544	6	2	8	531	511	4	2	9	252	97	
66	13	6	64	-49	11	4	7	73	-58	7	2	8	146	119	-130	2	9	104	-483	
67	1	7	72	-15	12	4	7	219	212	8	2	8	697	684	6	2	9	268	-277	
68	2	7	99	-125	13	4	7	-48	11	9	2	8	171	176	7	2	9	-66	21	
69	3	7	136	113	1	5	7	821	831	10	2	8	475	477	6	2	9	292	-294	
70	4	6	-46	7	2	5	7	187	207	11	2	8	478	476	0	2	9	165	107	
71	5	6	144	154	3	5	7	265	267	12	2	8	322	316	10	2	9	224	226	
72	6	6	-45	83	4	5	7	286	280	13	2	8	451	454	7	2	9	224	226	
73	7	6	208	267	9	5	7	879	877	1	3	8	297	297	12	2	9	150	-159	
74	8	6	122	139	6	5	7	174	182	2	3	8	177	-179	1	3	9	761	-758	
75	9	6	160	162	7	5	7	528	626	3	3	8	399	394	2	3	9	80	-69	
76	10	6	94	52	8	5	7	76	32	4	3	8	212	-229	3	3	9	570	-578	
77	11	6	205	217	9	5	7	492	507	5	3	8	545	545	4	3	9	76	-67	
78	12	6	-47	50	45	6	3	8	-67	46	5	3	8	595	-598	1	4	9	-66	-31
79	13	6	749	-711	7	2	5	294	292	2	4	8	162	-166	4	3	9	498	-499	
80	14	6	174	-166	12	5	7	-50	50	8	3	8	132	116	7	2	9	270	119	
81	15	6	558	423	13	5	7	165	179	9	3	8	446	446	5	3	9	101	-97	
82	16	6	204	-226	0	6	7	1027	987	10	3	8	127	127	5	3	9	216	-225	
83	17	6	409	411	1	6	7	166	162	11	3	8	284	292	10	3	9	108	-112	
84	18	6	243	-256	2	6	7	952	956	12	3	8	75	83	11	3	9	144	-146	
85	19	6	66	-178	1	7	7	751	771	0	4	8	693	688	-142	8	5	10	199	
86	20	6	205	205	7	7	7	159	159	1	4	8	140	132	10	4	9	209	-207	
87	21	6	100	-100	2	7	7	158</												

Table 4. Continued.

H	K	L	F(0)	F(C)	H	K	L	F(0)	F(C)	H	K	L	F(0)	F(C)	H	K	L	F(0)	F(C)
6	6	1C	93	-89	5	1	11	271	279	4	3	11	94	-98	5	5	11	270	281
7	6	1C	124	6	1	11	-56	16	5	3	11	273	226	6	5	11	107	105	
1	7	0	-55	2	7	11	256	271	6	3	11	-44	0	6	11	354	354		
2	7	1C	88	57	8	1	11	86	86	7	3	11	175	188	6	6	11	-51	7
3	7	1C	-13	t3	9	1	11	106	118	8	3	11	-45	-39	2	6	11	337	335
4	7	1C	143	147	0	2	11	96	69	0	4	11	486	473	3	6	11	-49	18
5	7	1C	115	119	1	2	11	-65	-31	1	4	11	62	63	4	6	11	327	339
6	7	1C	104	120	2	2	11	145	134	2	4	11	430	437	1	7	11	132	122
6	8	1C	85	62	3	2	11	-61	-38	3	4	11	69	57	0	0	12	473	414
1	8	1C	-50	54	4	2	11	291	253	4	4	11	224	226	1	0	12	71	-82
2	8	1C	13	100	5	2	11	-57	-57	5	4	11	-23	-23	0	0	12	428	429
3	8	1C	-47	36	6	1	11	297	306	6	3	11	189	196	3	0	12	248	-52
4	8	1C	195	215	7	2	11	-51	34	7	4	11	62	-75	4	0	12	385	304
1	1	11	192	184	8	2	11	220	230	1	5	11	542	556	5	0	12	232	-249
2	1	11	-64	14	1	3	11	126	121	2	5	11	-55	38	6	0	12	320	347
3	1	11	274	267	2	3	11	101	-107	3	5	11	458	475	1	1	12	335	363
4	1	11	-61	10	3	3	11	149	143	4	5	11	74	91	2	1	12	-53	-21
															3	4	12	197	211
															4	4	12	118	-116

Table 5. Bond lengths (\AA) and angles ($^\circ$). Standard deviations in parentheses. A prime denotes an atom generated by the twofold axis of symmetry.

Te - S(2) = 2.380 (2)
 S(1) - S(2) = 2.080 (2)
 S(1) - O(1) = 1.433 (5)
 S(1) - O(2) = 1.434 (5)
 S(1) - C(1) = 1.779 (6)
 C(1) - C(2) = 1.401 (9)
 C(2) - C(3) = 1.379 (11)
 C(3) - C(4) = 1.388 (11)
 C(4) - C(5) = 1.400 (10)
 C(5) - C(6) = 1.378 (10)
 C(6) - C(1) = 1.390 (9)

$\angle \text{S}(2) - \text{Te} - \text{S}(2') = 97.61$ (6)
 $\angle \text{Te} - \text{S}(2) - \text{S}(1) = 103.46$ (7)
 $\angle \text{S}(2) - \text{S}(1) - \text{O}(1) = 103.1$ (3)
 $\angle \text{S}(2) - \text{S}(1) - \text{O}(2) = 110.0$ (3)
 $\angle \text{S}(2) - \text{S}(1) - \text{C}(1) = 104.5$ (3)
 $\angle \text{O}(1) - \text{S}(1) - \text{O}(2) = 119.9$ (3)
 $\angle \text{O}(1) - \text{S}(1) - \text{C}(1) = 109.6$ (3)
 $\angle \text{O}(2) - \text{S}(1) - \text{C}(1) = 108.6$ (3)
 $\angle \text{S}(1) - \text{C}(1) - \text{C}(2) = 119.0$ (5)
 $\angle \text{S}(1) - \text{C}(1) - \text{C}(6) = 117.2$ (5)
 $\angle \text{C}(1) - \text{C}(2) - \text{C}(3) = 117.6$ (6)
 $\angle \text{C}(2) - \text{C}(3) - \text{C}(4) = 120.0$ (7)
 $\angle \text{C}(3) - \text{C}(4) - \text{C}(5) = 121.0$ (8)
 $\angle \text{C}(4) - \text{C}(5) - \text{C}(6) = 120.5$ (7)
 $\angle \text{C}(5) - \text{C}(6) - \text{C}(1) = 117.1$ (6)
 $\angle \text{C}(6) - \text{C}(1) - \text{C}(2) = 123.8$ (6)

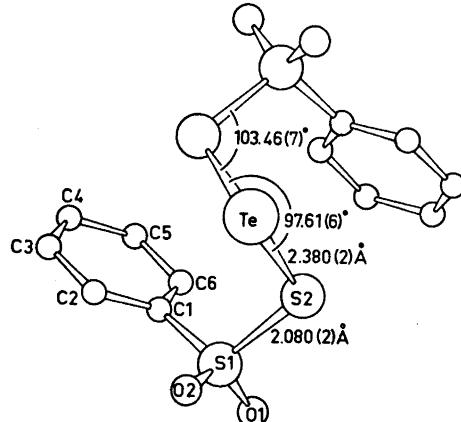


Fig. 1. The tellurium dibenzenethiosulphonate molecule as seen along the twofold rotation axis.

Since the molecule possesses a two-fold axis of symmetry, the non-planar S—S—Te—S—S chain occurs in the *trans* form, with the terminal sulphur atoms on opposite sides of the plane through the three middle atoms.

The greatest difference between the two investigations is in the Te—S(2) bond length, which is now found to be 2.380(2) Å. The older value is 2.41 Å, with a probable error of 0.03–0.04 Å. It is interesting to note that a Fourier map from the earlier study indicates a bond length of 2.39 Å.

The dihedral angle between the S(1)—S(2)—Te and the S(2)—Te—S(2') planes is 97.4°. The S(1)—S(2) bond length, 2.080(2) Å, is in excellent agreement with the earlier value. In the complex, *trans*-dibzenethiosulphonato-bis(trimethylenethiourea)tellurium(II),¹ the S—S bond length is 2.018(4) Å. This difference in the S—S bond lengths is thought to be caused by the difference in the nature of the Te—S bond in the complex relative to the uncomplexed compound, which is the subject of the present study.

The coordination around the S(1) atom is approximately tetrahedral, with the greatest deviation for the O(1)—S(1)—O(2) angle, which is 119.9(3)°. The bond lengths and angles involved in this coordination are in good agreement with the values given by Øyum and Foss.²

The atoms of a least squares plane through the benzene ring do not deviate more than 0.013 Å from the plane. S(1) is 0.095 Å from this plane, on the same side as Te, while O(1) and O(2) are 0.779 and 0.012 Å from it, on the opposite side.

Excluding the S(1)—S(2) bond length, the dimensions of the benzenethiosulphonate group are very similar to the dimensions found for the same group in the complex referred to above. In both compounds one of the oxygen atoms is near the plane of the benzene ring, and the C(6)—C(1)—C(2) angle is somewhat large.

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