mediate into tetracyclines and the failure to introduce methyl groups into resaceto-phenone for the formation of clavatol in Aspergillus clavatus. Recently Steward and Packter reported similar conclusions from their studies on gloriosin formation in Gliocladium roseum.

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Hydrothermal Preparation of Tellurium Compounds
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The tellurates $\text{M}_2\text{TeO}_6$ containing the metals Cu, Zn, and Hg, the silver tellurate, $\text{Ag}_2\text{TeO}_4$, and the lead tellurate, $\text{Pb}_2\text{H}_2\text{TeO}_4$, were prepared by Jander and Kienbaum using precipitation from aqueous solutions. The precipitation of the compounds is mostly initiated by the formation of amorphous tellurates of varying compositions, which on aging yielded crystalline products. The tellurates $\text{M}_2\text{TeO}_6$ containing the metals Mg, Mn, Ni, and Cu were prepared by Bayer using solid-state reactions from corresponding reagent-grade oxides. The tellurates $\text{M}_2\text{TeO}_6$ of Sc, In, Y, and the rare-earth elements were prepared by Natansohn by reacting the sesqui-oxides and orthotelluric acid at elevated temperatures.

Strontium tellurite, $\text{SrTeO}_3$, $\text{H}_2\text{O}$, was prepared by Ivankova, Samplavakaya, and Karapet'anya by precipitating a sodium tellurite solution with a strontium nitrate solution. The product obtained showed an endothermic effect in the temperature range 250—400°C, associated with the loss of water.

Only a limited number of tellurium compounds have been prepared by using hydrothermal technique. Compounds of the composition $\text{MTeO}_3(\text{OH})_n$ and $\text{MTeO}_4\text{OH}$ containing K and Rb were hydrothermally prepared by Lammers. Two modifications of tellurium oxide, $\text{TeO}_2$, and polymeta-telluric acid ($\text{TeO}_3\text{H}_2\text{O})_n$ were prepared by Maurin and Moret using hydrothermal technique.

We wish to report the hydrothermal synthesis of the tellurium compounds listed in Table 1.

Expts. Nos. 1—5. A solution of 0.01 M orthotelluric acid and solutions of 0.03 M of the respective metal ions were used. 20 ml of the metal ion solution was mixed with 20 ml of the orthotelluric acid solution in a silver ampoule and the precipitate was heated with the mother liquid in a 100 ml pressure bomb, as indicated in Table 1. The crystalline prod-

Corrigendum to "4-Methylsulphonylbutylglucosinolate Ion, the Natural Thiogluco side Precursor of Ery sosin"*
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Through a regrettable oversight, formula (II), p. 2875 is erroneously drawn: the hydroxy-substituents at C-2, C-3, and C-4 in the pyranose-ring should be reversed.

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Table 1. Experimental conditions for hydrothermal preparation of tellurium compounds.

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<td>2</td>
<td>Zn₅TeO₆</td>
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<tr>
<td>4</td>
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<td>6</td>
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<td>3400</td>
<td>73</td>
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<tr>
<td>7</td>
<td>Y₅TeO₆</td>
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<tr>
<td>11</td>
<td>SrTeO₅H₂O</td>
<td>180</td>
<td>10</td>
<td>40</td>
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</table>

Products were washed with water, dried at 100°C, and the X-ray powder patterns were obtained with a de Wolff-Guinier camera.

Expts. Nos. 6—10. Sealed gold tubes with a volume of 2 ml and 7 ml pressure bombs were used. In a typical preparation the reaction of 0.1 g of a metal sesquioxide-orthotelluric acid mixture in a 1:1 mole ratio in 0.7 ml of water yielded the orthotellurate at the experimental conditions listed in Table 1. In the experiments Nos. 6 and 9, crystalline In(OH)₃ and EuOOH were used and the metal compound-orthotelluric acid mole ratio was 2:1. The crystalline products were treated and the powder patterns were recorded as described for Expts. Nos. 1—5. The compounds prepared in Expts. Nos. 1—10 were all identified by the X-ray powder patterns.

Expt. No. 11. 20 ml of a 20% strontium nitrate solution was mixed with 20 ml of a 10% sodium tellurite solution in a silver ampoule and the precipitate was heated with the mother liquid in a 100 ml pressure bomb as indicated in Table 1. The crystalline product was washed with water and dried at 25°C. The DTA diagram was obtained with a Du Pont 900 Differential Thermal Analyzer over the temperature range from 25°C to 500°C using a heating rate of 20°C/min. A phase transformation was observed at 345°C. This is in agreement with the results reported in Ref. 4. The loss of weight on heating to 450°C for 6 h was determined. (Found: loss (H₂O) 5.4. Calc. for SrTeO₅H₂O: H₂O 6.4.) The infrared spectrum was recorded over the frequency range 300—4000 cm⁻¹ on a Beckman IR10 spectrophotometer using a pellet of a mixture of 3 mg sample in 300 mg CaF. The infrared spectrum had a broad band from 2700 cm⁻¹ to 3600 cm⁻¹ characteristic for a water-containing compound. The Guinier powder pattern of the sample and of SrTeO₅H₂O precipitated as reported in Ref. 4 were identical.

The investigation shows that amorphous and ultramicrocrystalline tellurium compounds can be converted to crystalline compounds using hydrothermal technique and that tellurates can be prepared at temperatures much lower than the temperatures necessary for solid-state reactions.

Acknowledgements. We are indebted to Professor S. Brodersen of the Department of Chemical Physics, Aarhus University, for the use of the spectrophotometer, to Professor L. T. Muus of the Department of Physical Chemistry, Aarhus University, for the use of the DTA equipment, and to Professor S. E. Rasmussen for his interest in this work.


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