from the latter solvent, as the dichloride dihydrate.

The salt forms plates {100} extended along the b axis in most cases, and bounded by {001}. There is perfect cleavage along the a plane, and a persistent tendency of twinning on this plane, but un-twinned crystals were found in a crop from a mother liquor. The unit cell dimensions are, $a = 11.96$ Å, $b = 11.48$ Å, $c = 17.08$ Å, $\beta = 112\frac{1}{2}°$, and there are four formula units per unit cell; density, calc. 1.94, found 1.93 g/cm³. The space group, from systematic absences, is $C_{2h}^5 - P2_1/c$. The reflections are very weak when $k + l$ is odd (particularly evident on oscillation photographs about the bc diagonal) which is typical for heavy atoms in centres of symmetry in this space group; however, such molecular symmetry is very improbable for stereochemical reasons. The same intensity distribution would result from tellurium atoms in general, fourfold positions x,y,z with x and $z = 0$ or $y = 0$.

1. Foss, O. and Hauge, S. *Acta Chem. Scand.* **13** (1959) 1252; **15** (1961) 1616.
2. Foss, O. and Hauge, S. *Acta Chem. Scand.* **15** (1961) 1623.
3. Freedman, L. D. and Corwin, A. H. *J. Biol. Chem.* **181** (1949) 601.
4. Johnson, T. B. and Edens, C. O. *J. Am. Chem. Soc.* **64** (1942) 2706.
5. Chase, B. H. and Walker, J. *J. Chem. Soc.* **1955** 4443.
6. Foss, O. and Fossen, S. *Acta Chem. Scand.* **15** (1961) 1618.
7. Foss, O. and Hauge, S. *Acta Chem. Scand.* **15** (1961) 1615.
8. *International Tables for X-ray Crystallography*, Vol. I, The Kynoch Press, Birmingham 1952, p. 101.

Received October 8, 1961.

Correction to "A Miniature Solubility Column and its Application to a Study of the Solubility of Red Mercury(II) Oxide in Acid 3 M $NaClO_4$ Solutions" *

DAVID DYRSSEN and VALENTINE TYRRELL

Department of Inorganic Chemistry, Royal Institute of Technology, Stockholm 70, Sweden

A systematic error in the $-\log[H^+]$ values quoted in this paper, arising from an arithmetical error, has been detected. Throughout they should be reduced by 0.142 units, *e.g.*, all points in Fig. 2 should be shifted 0.142 units to the left and $-\log[H^+]$ at the intersection of the asymptotes in Fig. 2 should be 3.08 not 3.22. As a result, the published equilibrium constants must be altered as follows:

$$\log {}^*K_1 + \log {}^*K_2 = -6.16 \pm 0.08$$

$$\log {}^*K_{s0} = 2.41 \pm 0.10$$

This new value of $(\log {}^*K_1 + \log {}^*K_2)$ agrees with Ahlberg's potentiometric value [1] within the expected limits of error. The values of $\log {}^*K_1 - \log {}^*K_2 = -0.30 \pm 0.30$ and $\log K_{s2} = -3.75 \pm 0.01$ remain unchanged.

In the calculation of the solubility product (pK_{s0}) of $Hg(OH)_2$, Lagerström's value [2] of $pK_w = 14.03$ was used. This is applicable to 3 *molal* solution of sodium perchlorate; the measured value for the 3 molar solutions used in the solubility work is 14.22 [3]. Using this and the revised value of $\log {}^*K_{s0}$, pK_{s0} becomes 26.0.

1. Ahlberg, I. *To be published.*
2. Lagerström, G. *Acta Chem. Scand.* **13** (1959) 722.
3. Ingri, N., Lagerström, G., Frydman, M. and Sillén, L. G. *Acta Chem. Scand.* **11** (1957) 1039.

Received October 12, 1961.

* *Acta Chem. Scand.* **15** (1961) 393.