

experiments where $^{15}\text{NH}_4\text{Cl}$, ATP, MgCl_2 , acetyl glutamate and aspartate- ^{14}C were substrates. The influence of increasing amounts of ^{13}C on the isotopic pattern of the formed USA was studied. The results were in full agreement with the scheme, since the ^{15}N -content of USA sharply decreased with increasing amounts of CAP, while the ^{13}C -content increased and finally reached the same level as the ^{14}C -content. Similar results were obtained with labeled citrulline in place of ammonia.

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Received June 22, 1955.

The Preparation of 2,6-Dichlorophenylacetic Acid

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Ortho halogen substituted benzoic, phenylacetic and phenoxyacetic acids have in recent years been frequently tested as to their auxin activity, among others by Thimann¹. The activity of a benzoic acid is greatly increased by di-substitution of halogen atoms in 2,6-position, while the same substitution in a phenoxyacetic acid completely destroys the activity.

The 2,4,6-trichlorophenylacetic acid was recently prepared by Jönsson², but the 2,6-dichlorophenylacetic acid is not earlier described. By a method related by Fierz-David and Blangey³ 2,6-dichlorotoluene was prepared. This compound was chlorinated to benzyl chloride which was treated with potassium cyanide to benzyl cyanide and hydrolysed to phenylacetic acid as Staedel⁴ has described.

For the sake of control a sample of the 2,6-dichlorotoluene was oxidized, yielding pure 2,6-dichlorobenzoic acid with m. p. 140–141°.

Experimental: 2,6-Dichlorobenzyl chloride. 16.1 g of 2,6-dichlorotoluene was chlorinated with dry chlorine at 180° in ultra-violet rays until an increase of the weight of 3.5 g had been obtained. The product was fractionated *in vacuo* yielding 12 g (61 %) of colourless oil, boiling at 114–119°/13 mm, melting point 11–12°.

2,6-Dichlorobenzyl cyanide. 6.5 g of benzyl chloride was refluxed 5 hours with 2.7 g of potassium cyanide and 30 ml alcohol. The alcohol was distilled off from the solid.

2,6-Dichlorophenylacetic acid. The before mentioned cyanide was refluxed with 1 N sodium hydroxide overnight. After extraction with ether the mixture was acidified with 2 N hydrochloric acid. The dichlorophenylacetic acid was obtained as a colourless solid (4.0 g; 60 %) which was repeatedly crystallised from aqueous ethanol. Melting point 158–159°. (Found: Equiv. wt. 205.4; Cl 34.41 %. Calc. for $\text{C}_8\text{H}_6\text{O}_2\text{Cl}_2$: Equiv. wt. 205.0; Cl 34.59 %).

The author wishes to express his thanks to Professor Arne Fredga for valuable discussions.

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Received June 21, 1955.

The Structure of Tellurium Dibenzenethiosulphonate

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This compound¹, $\text{Te}(\text{S}_2\text{O}_2\text{C}_6\text{H}_5)_2$, is an analogue of tellurium dimethanethiosulphonate, $\text{Te}(\text{S}_2\text{O}_2\text{CH}_3)_2$, and of salts of telluropentathionic acid, $\text{Te}(\text{S}_2\text{O}_5\text{OH})_2$. The crystal structure has been worked out and found to be closely analogous to that of